

### Heats of formation (298 K) of atmospheric gases

	$\Delta H_f^\circ (\text{kJ mol}^{-1})$
$\text{O}_{2(g)}$	0
$\text{O}_2^*_{(g)}$	95
$\text{O}_{3(g)}$	144
$\text{O}_{(g)}$	249
$\text{O}^*_{(g)}$	440
$\text{N}_{2(g)}$	0
$\text{N}^*_{(g)}$	473
$\text{NO}^*_{(g)}$	90.4
$\text{NO}_2^*_{(g)}$	33.3
$\text{NO}_3^*_{(g)}$	72.8
$\text{N}_2\text{O}_{(g)}$	104
$\text{NOCl}_{(g)}$	53.6
$\text{HNO}_{3(g)}$	-135
$\text{Cl}_{2(g)}$	0
$\text{Cl}^*_{(g)}$	106
$\text{ClO}^*_{(g)}$	102
$\text{ClO}_{2(g)}$	121
$\text{HCl}_{(g)}$	-95.4
$\text{H}^*_{(g)}$	218
$\text{OH}^*_{(g)}$	39
$\text{H}_2\text{O}_{2(g)}$	-136
$\text{H}_2\text{O}_{(g)}$	-242
$\text{CH}_3\text{Cl}_{(g)}$	-82
$\text{CH}_3^*_{(g)}$	146
$\text{HOCl}_{(g)}$	-79.5
$\text{ClONO}_{2(g)}$	26.4